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2 Section A Answer all questions. Write your answers in the boxes provided. 1. Ethanedioic acid is a diprotic acid. A student determined the value of  $x$  in the formula of hydrated ethanedioic acid,  $\text{HOOC}(\text{COOH})_x\text{H}_2\text{O}$ , by titrating a known mass of the acid with a  $0.100 \text{ mol dm}^{-3}$  solution of  $\text{NaOH}(\text{aq})$ .  $0.795 \text{ g}$  of ethanedioic acid was dissolved in distilled water and made up to a total ...

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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.  $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$   $3 O_2 + 2 CO_2(g) + 3 H_2O$   $1411 \text{ kJ mol}^{-1}$   $286 \text{ kJ mol}^{-1}$   $1560 \text{ kJ mol}^{-1}$  (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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